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## Grade 10 chemistry practice exam with answers

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MCQ: Na<sub>2</sub>SO<sub>4</sub> is a/an salt common salt normal salt MCQ: KHCO<sub>3</sub> is a/an common salt salt normal salt salt MCQ: What is the term used for a salt that is formed by partial neutralization of a polyhydric base? Common salt common salt normal salt salt MCQ: NaHSO<sub>4</sub> is a/an salt base common salt normal salt salt MCQ: What is the term used for a salt containing a replaceable OH group? Common salt common salt salt normal salt You can create printable tests and worksheets from these grade 10 chemistry questions! Select one or more questions using the check boxes above each question. Then click the Add Selected Questions button to a test before moving to another page. Previous page 1 of 84 Next previous page 1 of 84 Questions next 14 Total attempts: 20734 6th Grade 7th Grade 8th Grade 8th Student Light high school from a glow stick is a sign of a chemical change? Are ion bonds formed when frontally charged molecules are attracted to each other? In an anion, is the number of protons equal to the number of electrons? Is a supplier label for a dangerous chemical required by law to have a dotted border? Is the mass number of an atom the sum of its protons and neutrons? The substance is a liquid at 25 C The substance has a density of 2.50 g/cm<sup>3</sup> The substance burns brilliantly in the air The substance will not dissolve in water What element would have atoms with the same number of external electrons as carbon atoms? Which would be made up of molecules? Of the following, which would be the best electricity conductor? What is the correct chemical formula for compound aluminum chloride? What is the correct chemical formula for compound copper fluoride (II)? What is the correct name for the compound K<sub>2</sub>O? ©2020 Walmart Stores, Inc. 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Quizzes & Practice Tests Summary Chapter 1: Acids Bases and Salts MCQ Chapter 2: Biochemistry MCQs Chapter 3: Characteristics of Acids Bases and Salts MCQ Chapter 4: Chemical Equilibrium MCQs Chapter 5: Chemical Industries MCQ Chapter 6: Environmental Chemistry I Atmosphere MCQ Chapter 7: Environmental Chemistry II Water MCQs Chapter 8: Hydrocarbons MCQ Chapter 9 : Organic Chemistry MCQs Chapter 10 : Atmosphere MCQs Key Response Chapter 1 Acids Bases and Salts MCQ 1: When an acid reacts with metal carbonate, the products are A. salt B. water C. carbon dioxide D. all upper MCQ 2: Preservatives are used to store A. foods B. acids C. bases D. water MCQ 3: Phenolphthalein in acid solution is A. colorless B. pink color C. yellow yellow D. colored orange MCQ 4 : The process in which acids (H) and bases (OH-) react to form salts and water is called A. Neutralization B. hydrogen C. halogenation D. sublimation MCQ 5: A substance that gives a pair of electrons to form a coordinated covalent bond is called A. Lewis acid B. Lewis base C. Bronsted-Lowry acid D. Bronsted-Lowry base MCQ 6: Corrosive effect on the skin is caused by A. acids B. bases C. water D. : A species that can accept a proton is called A. acid B. base C. neutral compound D. cation MCQ 8: If the value of the pH is greater than 7, so the solution is A. acidic B. basic C. neutral D. salesa MCQ 9: The pH of the water is A. 8 B. 3 C. 2 D. 7 MCQ 10: The amount of hydrochloric acid (HCl) secreted daily by the gastric glands is A. 1 on B. 3 on C. 2 illuminated D. 4 ACCQ 11 illuminated: KOH is used in the manufacture of A. Discharge detergent B. antacid C. cement D. liquid soap MCQ 12: The acid used in the manufacture of fertilizers and explosives is A. nitric acid B. sulphuric acid C. phosphoric acid D. hydrochloric acid MCQ 13: As an electrolyte, the water is A. strong B. neutral C. D weak : When a metal replaces the hydrogen atom , the compound form is A. oxide B. ether C. salt D. alcohol MCQ 15: The pH in which methyl red changes color is A. 7 B. 5.5 C. 3.8 D. 9 MCQ 16: In pure water, concentrations of A. H and OH- ions are equal B. ions H, H : Substances reacting with both acids and bases are called A. neutral B. conjugate bases C. amphoteric substances D. conjugate MCQ 18 acids : The pH (hydrogen power) value of black coffee is A. 7 B. 8 C. 3 D. 5 MCQ 19: If the concentration of H itself is greater than 1 x 10<sup>-7</sup> then the solution is A. neutral B. base C. Q: Bronsted-Lowry acid in the H<sub>2</sub>O H<sub>2</sub>O reaction is A. H<sub>2</sub>O B. NH<sub>3</sub> C. OH- D. NH<sub>4</sub> MCQ 21: Acids ionize in water to produce A. OH- ions B. ions H + C. SO<sub>4</sub>-2 ions D. H<sub>2</sub>O molecules MCQ 22: Salt among the following is A. HCl B. KCl C. HNO<sub>3</sub> D. H<sub>3</sub>4 MCQ 23: The bases leave the blue sunflower A. red B. pint C. black D. unchanged Chapter 2 Biochemistry MCQ MCQ 1 : The 10 consists of A. ribose sugar B. phosphate unit C. nitrogen base D. all higher than MCQ 2: Which of the following activates more than 100 different enzymes? A. vitamin A B. vitamin B. vitamin C. vitamin D. vitamin D MCQ 3: Which of the following are obtained from fruits, vegetables and cereals? A. monosaccharides B. sucrose C. cellulose D. starch MCQ 4: Which of the following is not included in fat-soluble vitamins? A. vitamin A B. vitamin D C. vitamin E D. vitamin B MCQ 5: Which of the following is used to make the spirit rectified by the fermentation process? A. cellulose B. starch C. glucose D. fructose MCQ 6: The molecular formula of fructose is A. C<sub>12</sub>H<sub>22</sub>O<sub>11</sub> B. C<sub>18</sub>H<sub>32</sub>O<sub>16</sub> C. C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> D. C<sub>7</sub>H<sub>14</sub>O<sub>7</sub> MCQ 7: In the human diet, all cholesterol comes from A. animal products B. plant products C. minerals D. beans MCQ 8: The general formula of Monosaccharides is A. C<sub>n</sub>H<sub>2n</sub> 2 B. C<sub>n</sub>H<sub>2n</sub> C. C<sub>n</sub>H<sub>2n</sub>-2 D. (CH<sub>2</sub>)<sub>n</sub> MCQ 9 : The condensed structural formula of palmitic acid is A. CH<sub>3</sub>-CH<sub>2</sub>-CH<sub>2</sub>-COOH B. CH<sub>3</sub>-(CH<sub>2</sub>)<sub>4</sub>-COOH C. CH<sub>3</sub>-(CH<sub>2</sub>)<sub>14</sub>-COOH D. CH<sub>3</sub>-(CH<sub>2</sub>)<sub>16</sub>-COOH MCQ 10: How many amino acids are synthesized by our bodies? A. 10 B. 20 C. 30 D. 40 MCQ 11: The condensed structural formula of caproic acid is A. CH<sub>3</sub>-CH<sub>2</sub>-CH<sub>2</sub>-COOH B. CH<sub>3</sub>-(CH<sub>2</sub>)<sub>4</sub>-COOH C. CH<sub>3</sub>-(CH<sub>2</sub>)<sub>14</sub>-COOH D. CH<sub>3</sub>-(CH<sub>2</sub>)<sub>16</sub>-COOH MCQ 12: Vitamin B is necessary for A. eyes and skin B. Energy production in cells C. healing wounds and preventing colds D. bones and teeth MCQ 13 : Is starch an example of? A. monosaccharides B. oligosaccharides C. polysaccharides D. lipids MCQ 14: Which of the following causes of dry skin when taken in excess? A. vitamin A. vitamin B. vitamin B. vitamin C. vitamin C D. vitamin D MCQ 15: Vitamin Deficiency A cause A. B. anemia and bleeding gums C. scurvy D. rickets and Osteomalacia MCQ 16: Vitamin B is soluble in A. fat B. water C. alcohol D. ether MCQ 17: Which of the following is the key to DNA's ability to store genetic information and transmit it from generation to generation? A. Double stranded structure of DNA B. deoxyribose sugar C. phosphate unit D. nitrogen base MCQ 18: Which of the following is obtained by heating bones and tendons in water? A. gelatin B. enzyme C. amylase D. lactase MCQ 19: Which of the following is a code that

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